



# Amitrakshar International Journal

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## STUDY OF SYNTHESIS, SPECTROSCOPIC CHARACTERIZATION AND BIOLOGICAL EVALUATION OF NI(II), CU(II) AND ZN(II) COMPLEXES OF DIPHENYLDITHIOCARBAMATE

BY

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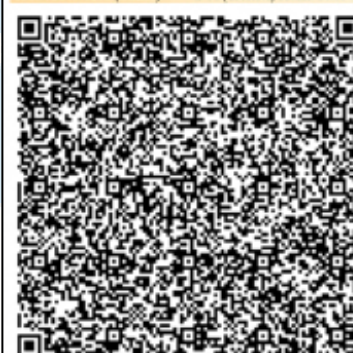
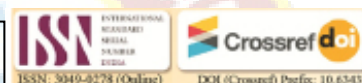
Guided by

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### Abstract:

Ni(II), Cu(II) and Zn(II) complexes of diphenyldithiocarbamate have been synthesized with the aim of developing antioxidants, potential antibacterial and antifungal agents. Synthesized metal complexes were characterized by magnetic susceptibility, elemental analysis; conductivity measurements, mass spectra, infrared, ultra violet/visible and nuclear magnetic resonance verified the chemical composition of products. Magnetic susceptibility and spectroscopic studies suggested the predicted geometry of the diphenyldithiocarbamate complexes as octahedral geometry of Ni(II) complex distortion square-planer geometry distorted towards tetrahedral for Cu(II) and Zn(II) complexes. Synthesized metal complexes exhibited remarkable antioxidant potential as compared to standard. Zn(II) complex of diphenyldithiocarbamate exhibited IC<sub>50</sub> of 31.45±0.31 μM while the standard, butylated hydroxytoluene exhibited IC<sub>50</sub> of 44.67±0.45 μM antioxidant activity. Synthesized metal complexes showed very good antibacterial and antifungal potential against *Rhodococcus*, *Actinomyces viscosus*, *Bacillus subtilis*, *Escherichia coli* and the fungi, *Aspergillus niger*, *Aspergillus flavus*, *Candida* and *Acetomyces*. It's concluded that synthesized metal complexes showed better antioxidant, antibacterial and antifungal potential as compared to diphenyldithiocarbamate as well as the standard and might turn out to be a valuable treatment of numerous kinds of diseases.

**Keywords:** *Diphenyldithiocarbamate (DPDTC), spectroscopic studies, antioxidant, antibacterial, antifungal activities*



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### Introduction

Dithiocarbamates (DTCs) are the synthetic compounds that involved in the forming of complexes with most of the stabilize transition metals in a various oxidation states. DTCs draw much attention in recent years and a large number of DTCs complexes have been synthesized on a large scale in the last 40-50 y [1-4]. A new synthetic strategy has been reported to prepare the DTCs[5]. They also have a variety of applications in agriculture, like pesticides and fungicides, in therapy, against alcoholism as arrestors of human immune deficiency virus infections such as AIDS in fungal diseases as well as in the rubber industries because they have strong metal binding capacity[6-10]. Coordination compounds of DTCs have been widely studied because metal complexes of this have been reported as an antitumor, antibacterial and antifungal agent[11-13]. Previously reported studies on the preparation of DTCs obtained from amino acids and their coordination compounds with several transition metal cations have been reported[14-17]. Some metal complexes of DTCs have been also reported to show de-toxicant and immune pharmacological properties[18]. The

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DTC stoichiometry has much external consideration in recent years due to its ability to act as a bidentate ligand. DTCs such as diethyl DTC and pyrrolidine DTC (PDTC) display cytotoxic properties and they have been used to treat metal poisoning[4-6]. The complexes are thermodynamically stable, so they are used as chelating agent for the extraction of trace metals[5,7]. Metal complexes of DTC present a wide range of applications in agriculture, medicine, industry, and analytical and organic chemistry]. The metal complexes of DTC with transition metals i.e. Zn, Ni, Cu, Mn, Pt show enhanced biological activities such as antibacterial, antimalarial, and fungicidal activities compared to free ligands due to drop in polarity of metals after complexation[9,10]. DTC metal complexes of Zn, Ni, and Cu, also showed remarkable antitumor properties[11]. Particularly, Ni(II) and palladium (II) play an important role in both fundamental research and application of their physical properties. Nickel and palladium complex with PDTC has been studied extensively. DTC ligands have been reported in the dithiocarboxy group as the ligator group and behave like a bidentate ligand in most of complexes. DTCs act as uni-negative bidentate ligands and they can also found in both hexa and tetra coordination geometry of the many transition metal complexes[19-25]. The present research work involved the synthesis of the ligand diphenyl DTC (DPDTC) and its respective metal complexes with Zn(II), Ni(II) and Cu(II) to enhance the antibacterial, antifungal and antioxidant activities of DPDTC by complexation. Further, all of complexes were characterized using different spectroscopic techniques. Antioxidant and antimicrobial activities of the complexes were also studied.

## MATERIALS AND METHODS

Carbon disulfide (CS<sub>2</sub>), diphenylamine, sodium hydroxide, ZnCl<sub>2</sub> .2H<sub>2</sub> O, CuCl<sub>2</sub> .2H<sub>2</sub> O and NiCl<sub>2</sub> .6H<sub>2</sub> O and other organic solvents were purchased from Merck and used without further purification. Elemental analyses of all compounds were performed on a Fisons elemental analyser. Conductance measurements of the synthesized complexes measured on conductivity meter, model WD-35604-00 at room temperature. Infrared (IR) spectra were recorded in Fourier-transform infrared spectroscopy (FTIR) spectrophotometer, Nicolet IS50, within the range of 400-4000 cm<sup>-1</sup>. Electronic spectra were recorded for solution of the ligand and their metal complexes in dimethylformamide as on Jasco UV/Vis spectrophotometer model V-770 UV/ Vis. Magnetic susceptibility of all the synthesized complexes was determined by using the Holmarc's magnetic susceptibility with model HO-ED-EM-07 at room temperature. 1 H-nuclear magnetic resonance (NMR) spectra and 13C-NMR were recorded in dimethyl sulfoxide (DMSO)-d<sub>6</sub> at room temperature using tetramethylsilane as internal standard on a Bruker 400 MHz and 101 MHz spectrophotometer. Mass spectra of all the complexes were performed on Finnigan MAT LCQTM mass spectrometer.

### Synthesis of ligand Na-DPDTC:

Na-DPDTC was synthesized using the previously reported methods with some modification[26]. DTCs ligand was prepared by mixing methanolic solution of diphenylamine (0.0845 g, 0.5 mmol), 5 ml cold 10 N solution of sodium hydroxide and 0.030 ml (0.5 mmol) CS<sub>2</sub> added drop-wise. The resulting mixture was kept at constant stirring for 3-4 h at temperature 273 k. The temperature maintained to avoid any possible decomposition by using a mixture of sodium salts, melting ice and water. Off-white precipitate of DTC ligand was filtered, washed with ethanol and dried over silica gel. Yield: 90 %, colour: off-white, melting point (m.p.); 189-192, anal. calc. for C<sub>13</sub>H<sub>10</sub>NS<sub>2</sub> Na (molecular weight, MW: 267.34 g/mol): C, 58.40; H, 3.77; N, 5.24; S, 23.99; Na, 8.60 % and found: C, 58.36; H, 3.72; N, 5.28; S, 23.95; Na, 8.64 %. Selected FTIR (KBr, v, cm<sup>-1</sup>): (NCS/CN): 1465; (CS<sub>2</sub>)<sub>sy</sub>: 994; (CS<sub>2</sub>)<sub>asy</sub>: 1156; (S-M): 423. Electronic spectra (λ<sub>max</sub> in nm): 267, 287.

### Synthesis of metal complexes of DPDTC:

Metal complexes of DPDTC were prepared by taking the ligand to metal molar ratio of 2:1. An aqueous solution of ZnCl<sub>2</sub> .2H<sub>2</sub> O (0.25 mmol, 0.043 g), CuCl<sub>2</sub> .2H<sub>2</sub> O (0.25 mmol, 0.042 g) and NiCl<sub>2</sub> .6H<sub>2</sub> O (0.25 mmol, 0.059 g) was mixed separately with 0.5 mmol (0.134 g) DTC ligand dissolved in methanol. The resulting mixture was reflux for 2 h and the colored precipitates of DPDTC metal complexes were obtained for each case and the solid precipitated was filtered, washed with diethyl ether-methanol (1:1 ratio) and dried over silica gel.

### Zn[S<sub>2</sub> CN(C<sub>6</sub> H<sub>5</sub>)<sub>2</sub>] complex:

Yield: 88 %, colour: purple, m.p.; 239-241, λ<sub>m</sub>: 342 Ohm<sup>-1</sup> cm<sup>2</sup> mol<sup>-1</sup>, μ<sub>eff</sub>: 3.67. Analysis: calculated for C<sub>26</sub>H<sub>20</sub>N<sub>2</sub> S<sub>4</sub> Zn (MW: 554.10 g/mol): C, 56.36; H, 3.64; N, 5.06; S, 23.23; Zn, 11.72 % and found: C, 56.31; H, 3.60; N, 5.09; S, 23.27; Zn, 11.68 %. Selected FTIR (KBr, v, cm<sup>-1</sup>): (NCS/CN): 1481; (CS<sub>2</sub>)<sub>sy</sub>: 987; (CS<sub>2</sub>)<sub>asy</sub>: 1128; (S-M): 431. Electronic spectra (λ<sub>max</sub> in nm): 259, 298, 345. 1 H-NMR (400 MHz, DMSO-d<sub>6</sub>) δ= 6.29- 7.56 (m, 20H, -C<sub>6</sub> H<sub>6</sub>). 13C-NMR (DMSO-d<sub>6</sub>) δ= 189.81 (-CS<sub>2</sub>), 141.92, 129.67, 127.55, 121.90 (-C<sub>6</sub> H<sub>5</sub>).

Cu(II)[S<sub>2</sub> CN(C<sub>6</sub> H<sub>5</sub>)<sub>2</sub>] complex: Yield: 83 %, colour: radish-brown, m.p.; 273-275, λ<sub>m</sub>: 653 Ohm<sup>-1</sup>cm<sup>2</sup> mol<sup>-1</sup>, μ<sub>eff</sub>: 1.81. Analysis: calculated for C<sub>26</sub>H<sub>20</sub>N<sub>2</sub> S<sub>4</sub> Cu (MW: 552.256 g/mol): C, 56.55; H, 3.65; N, 5.07; S, 23.22; Cu, 11.51 % and found: C, 56.60; H, 3.69; N, 5.02; S, 23.27; Cu, 11.54 %. Selected FTIR (KBr, v, cm<sup>-1</sup>): (NCS/CN): 1474; (CS<sub>2</sub>)<sub>sy</sub>: 978; (CS<sub>2</sub>)<sub>asy</sub>: 1131; (S-M): 427. Electronic spectra (λ<sub>max</sub> in nm): 379, 434, 547, 467. 1 H-NMR (400 MHz,



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DMSO-d<sub>6</sub>) δ= 6.17-7.60 (m, 20H, -C6 H<sub>6</sub>). <sup>13</sup>C-NMR (DMSO-d<sub>6</sub>) δ= 196.74 (-CS<sub>2</sub>), 142.75, 129.69, 127.25, 119.62 (-C6 H<sub>5</sub>)

### Ni(II)[S<sub>2</sub>CN(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>2</sub>] complex:

Yield: 79 %, colour: green, m.p.; 263-265, λ<sub>m</sub>: 557 Ohm-1cm<sup>2</sup> mol<sup>-1</sup>, μ<sub>eff</sub>: 3.67. Analysis: calculated for C<sub>26</sub>H<sub>24</sub>N<sub>2</sub>S<sub>4</sub>O<sub>2</sub>Ni (MW: 583.43 g/mol): C, 53.52; H, 4.15; N, 4.80; S, 21.98; Ni, 10.26 % and found: C, 53.48; H, 4.19; N, 4.79; S, 21.99; Ni, 10.23 %. Selected FTIR (KBr, v, cm<sup>-1</sup>): (NCS/CN): 1477; (CS<sub>2</sub>) sy: 996; (CS<sub>2</sub>) asy: 1143; (S-M): 429; (H<sub>2</sub>O): 2600-3600; (OH): 1648. Electronic spectra (λ<sub>max</sub> in nm): 367, 412, 670, 559. <sup>1</sup>H-NMR (DMSO-d<sub>6</sub>) δ= 6.21-7.52 (m, 20H, -C<sub>6</sub> H<sub>6</sub>), 4.41 (s, 2H, H<sub>2</sub>O). <sup>13</sup>C-NMR (400 MHz, DMSO-d<sub>6</sub>) δ= 201.42 (-CS<sub>2</sub>), 134.97, 123.53, 127.05, 124.94 (-C<sub>6</sub> H<sub>5</sub>).

### DPPH free radical scavenging activity:

The DPPH free radical controlling effect was concluded by using method proposed by Stasko et al.[27] To examine the activity of the Na-DPDTC and their metal complexes at various concentrations 1000, 500, 250, 150 μg/ml and 60 μg/ml and butylated hydroxytoluene (BHT) was used as a reference drug[28]. Then 3 ml solution of DPPH in methanol (0.01 g/250 ml methanol) was mixed with sample solution and mixed energetically and permitted to place in room temperature towards 1 h. The absorbance was measured at 517 nm against volatile flammable liquid alcohol as a void in spectrophotometer. All measurements were taken triplicate and lesser absorbance indicates high level of free radical controlling capacity[29,30]. Free radical scavenger effects of DPPH are estimated using the formula, DPPH inhibition =  $\frac{Abs_{control} - Abs_{sample}}{Abs_{control}} \times 100$ , where  $Abs_{control}$  is the absorbance of DPPH radical + DMSO and  $Abs_{sample}$  is the absorbance of DPPH radical + sample/standard.

### Antimicrobial activity:

Antimicrobial activity of Na-DPDTC and the synthesized metal complexes were tested using the agar-well diffusion method[31,32] against diverse bacteria, Rhodococcus, P. aeruginosa, B. subtilis and E. coli and fungi, A. niger, A. flavus, Candida and Acetomyces. The culture media were prepared by dissolving around 12 g nutrient agar in 600 ml distilled water for bacteria and 38 g of the potato dextrose agar in 1 l of distilled water was poured in it for the preparation of solution. The culture media were buffered at a pH of 7.4 and sterilized in autoclave at 121° for 15 min to purify the media[33,34]. All the bacterial and fungal strains were adjusted to 0.5 McFarland standard, which was prepared by adding 0.05 ml of 1 % w/v BaCl<sub>2</sub> · 2H<sub>2</sub>O in phosphate buffered saline (PBS) to 9.95 ml of 1 % v/v H<sub>2</sub>SO<sub>4</sub> in PBS, which is visually comparable to a microbial suspension containing approximately 1.5 × 10<sup>8</sup> CFU/ml. The stock solutions of synthesized complexes were aseptically filled (100 μl in each dish) into Petri dishes and then allowed to attain room temperature. After keeping the petri dishes in the flat position for one hour, the incubation period was allowed to proceed for 24 h at 37° for the bacterial cultures and for the fungal species 48 h at 37°[35], the development was checked with the assistance of zone reader in millimeter by measuring the zone of hindrance around the specimen well. Standard antifungal drug fluconazole and antibacterial drug ampicillin (100 μg/ml) were used for comparison under similar conditions. These studies were performed in triplicate, and the average was taken as the final reading.

## RESULTS AND DISCUSSION

Numerous methods for synthesis of DTCs were reported using the amine with CS<sub>2</sub>. In this investigation, synthesis of Na-DPDTC by the reaction of CS<sub>2</sub> with a diphenylamine in the presence of NaOH at 0° was reported. The temperature maintained to avoid any possible decomposition. The metal complexes of DPDTC were synthesized by reaction of ligand with metals salts and analytical data indicated the ligand to metals ratio 2:1 of all the synthesized complexes. The molecular formulae of the ligands and their metal complexes were proposed according to the data from spectral, elemental analysis, conductivity and magnetic susceptibility of metal complexes. The proposed mechanism for the synthesis of the compounds is shown in fig. 1.

Magnetic susceptibility of the Ni(II), Zn(II) and Cu(II) complexes of DPDTC have the magnetic moment in the range of 1.73-3.67 B.M. The magnetic moment of Ni(II) complex of DPDTC is 3.67 B.M, which agrees with the formation of octahedral complex. The Zn(II) and Cu(II) complexes of DPDTC usually have a distorted square planer configuration with magnetic moment values, 1.73 B.M and 1.81 B.M, respectively.

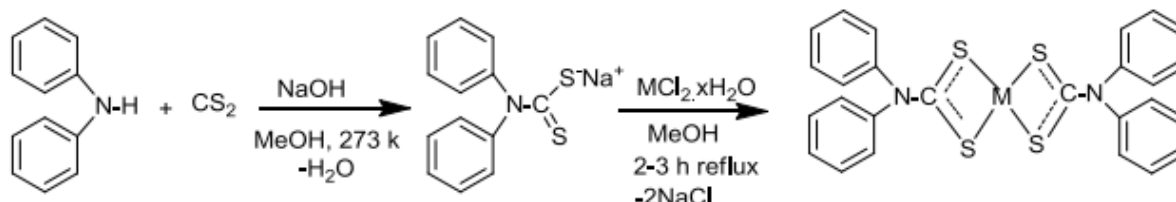


Fig. 1: Protocol for the synthesis of metal complexes of diphenyldithiocarbamate (DPDTC)



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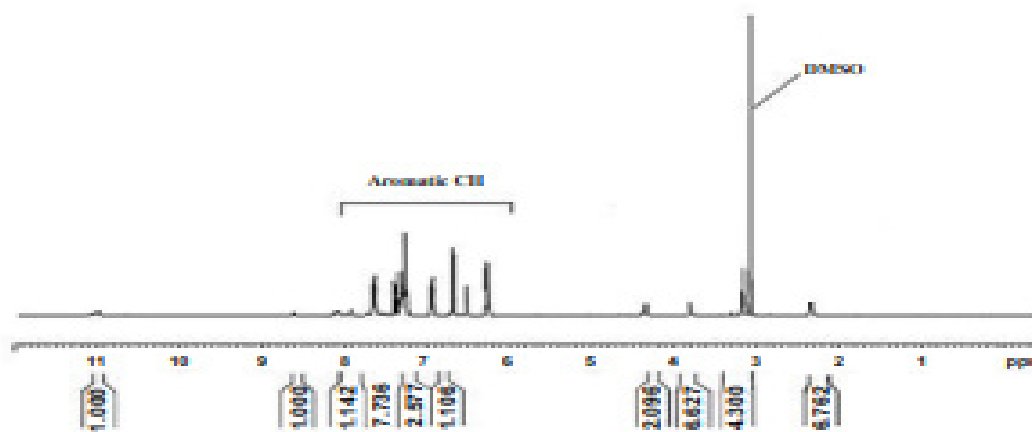
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Functional groups can be simply recognized in IR spectroscopy. Consequently, IR spectroscopy considered as valuable basis for attaining binding mode of the Na-DPDTC to metal ions in the complexes formation. The three main regions are of interest of the DPDTC and their metal complexes, which give the valuable information about the coordination behaviour of ligand to metals and have a strong agreement about structure of the complexes as in the region of 1456-1488  $\text{cm}^{-1}$  was assigned the stretching vibrations of  $\nu(\text{C-N})$  for  $\text{NCS}_2$  – with the shift of 9-16  $\text{cm}^{-1}$  in higher wave number in all the complexes, indicating its involvement in coordination with metal ions. IR spectra of all the synthesized complexes give the strong peak of asymmetric absorption frequency of  $\nu(\text{CS}_2)_{\text{asy}}$  at 1128-1156  $\text{cm}^{-1}$  with the shift of 13-28  $\text{cm}^{-1}$  to the lower wave number and symmetric absorption frequency of  $\nu(\text{CS}_2)_{\text{sy}}$  at 978-996  $\text{cm}^{-1}$  and which indicated that the bidentate ligand to the central metals. The sharp stretching vibration bands found in the spectra of the Ni(II) complex of DPDTC appear in medium peak at 3469  $\text{cm}^{-1}$  and weak peak at 1648, which are not observed in the spectra of the ligand and other complexes. New absorption bonds of M-S stretching mode of vibration are assigned within the range of 423-431  $\text{cm}^{-1}$  of all the synthesized complexes. These values are close agreement with the metal-sulphur donor ligands.

The UV/Vis spectra of the Na-DPDTC and its metal complexes have been recorded in DMSO at room temperature. UV spectra show the high absorption intensity due to presence of  $\text{NCS}_2$  chromosphere. In the spectrum of Na-DPDTC ligand two absorption bands were recorded at 287 and 267 nm, which are attributed to  $n \rightarrow \pi^*$  and  $\pi \rightarrow \pi^*$  transitions. The copper complexes of DPDTC exhibited two absorption bands in the visible region at 547 and 467 nm due to  $2 B_{1g} \rightarrow 2 A_{2g}$ ,  $2 B_{1g} \rightarrow 4 E_g$  transition, respectively[41]. A sharp band appeared at 556 nm is due to Jahn-Teller distortion, which indicate that the Cu(II) complex of DPDTC has the square-planar geometry. The nickel complex also exhibited the two absorption bands as  $3 A_{2g} \rightarrow 3 T_{1g}$ .

(P) and  $3 A_{2g} \rightarrow 3 T_{1g}$  (F) transitions in visible region at 670 and 559 nm, suggesting the octahedral geometry of the nickel complex[42]. The electronic spectra of the Zn(II) complex of DPDTC exhibited one broad peak at 345 nm, which assigned charge transfer of metal to ligand and d-d electronic transition was not exhibited due to the completely filled d-orbital. In general, Zn(II) complexes of DPDTC would have square-planar geometry due to four coordination bonds[43]. Strong bands appeared at 367 and 379 nm are assigned to transfer of charge from metal to ligand (M $\rightarrow$ LCT) and d-d bond for Ni(II) and Cu(II) complexes, respectively. Sulphur to metal charge transfer ( $L\pi \rightarrow MCT$ ) and d-d bond is assignable by the very intense bands at 412 and 434 nm for Ni(II) and Cu(II) complexes, respectively[44]. The  $^1\text{H}$  NMR spectra of synthesized metal complexes were recorded in DMSO- $d_6$ . These synthesized complexes appeared the multiples signals were in the region at  $\delta=7.60$ -6.29 ppm, which assigned to proton of phenyl group[45]. The  $^1\text{H}$  NMR spectra of Zn(II), Cu(II) and Ni(II) complexes are shown in fig. 2.  $^{13}\text{C}$  NMR spectra of the complexes were observed the week signal of  $\text{NCS}_2$  carbon atoms moieties at the region 189.81-201.42 ppm[46]. The signals of carbon due to aryl group were observed at the region 141.92- 121.90, 142.75-119.62 and 144.97-123.94 ppm in the Zn(II), Cu(II) and Ni(II) complexes, respectively. The  $^{13}\text{C}$  NMR spectra of Zn(II), Cu(II) and Ni(II) complexes are shown in fig. 3.

The mass spectra of the Na-DPDTC and all synthesized complexes have been recorded at 300° and 70 eV. All the spectra exhibit parent peaks due to molecular ions and proposed molecular formula of these complexes was confirmed by comparing their molecular formula weights with m/z values. The mass spectra of ligand show molecular ion peak (M $^+$ ) at m/z 267 resembles to the total molecular weight of the ligand. In the spectra of Zn(II), Cu(II) and Ni(II) shows the molecular ion peaks at m/z 554, m/z 552 and 583, respectively, supporting the composition of synthesized complexes. Mass spectra of Zn(II), Cu(II) and Ni(II) complexes are shown in fig. 4.





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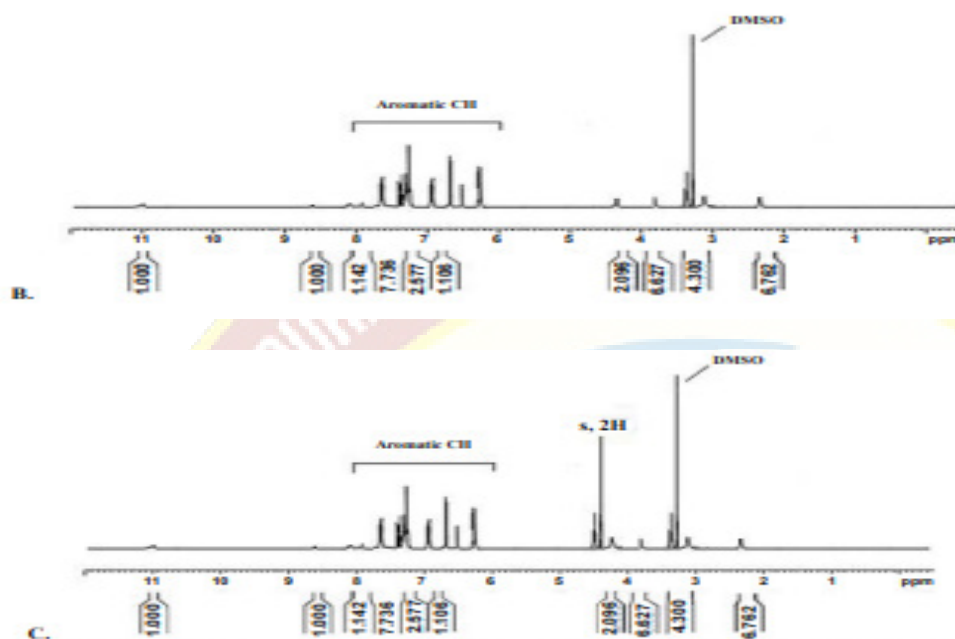
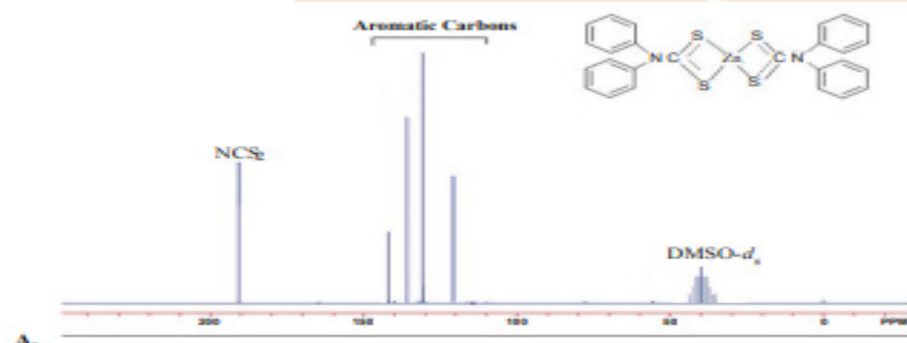


Fig. 2: <sup>1</sup>H NMR spectra of Zn(II)-, Cu(II)- and Ni(II)-DPDTC complexes

**<sup>1</sup>H NMR spectra of (A) Zn(II)-DPDTC complex; (B) Cu(II)-DPDTC complex and (C) Ni(II)-DPDTC complex**

DPPH assay is the most extensively used method for evaluating the oxidative potential of different metal complexes. Free radicals are generated by enzymatic and non-enzymatic processes and if the colours of these free radicals disappear, it is understood that the compounds have the ability to scavenge free radicals[47,48]. The antioxidant activity of ligand and synthesized complexes have been evaluated for DPPH radical scavenging assay against the BHT was used as a reference drug at various concentrations. All the complexes and ligand showed the good antioxidant activity. The inhibitory concentration (IC<sub>50</sub>) of Zn(II) was lower 31.45±0.31 μM, it also showed highest antioxidant activity than other compounds. The Ni(II) complex also exhibited highest antioxidant activity by showing lowest IC<sub>50</sub> value of 33.55±0.35 μM. But IC<sub>50</sub> value of Cu(II) of was lower 39.12±0.41 μM it also showed highest antioxidant activities than ligand as well as reference standard (BHT)[49]. The order of DPPH assay of the compounds is as Zn(II)>Ni(II)>Cu(II)>BHT>ligand. The oxidizing potential effects of the compounds are associated with in the presence of ligand and their complexes are exerting the actions by donation of hydrogen. IC<sub>50</sub> inhibition value of DPPH antioxidant activities of Na-DPDTC and synthesized complexes are presented in fig. 5.

The antibacterial activity of Na-DPDTC and their metal complexes were investigated against four bacterial strains i.e. two Gram-negative bacterial strains as *E. coli* and *P. aeruginosa* and two Gram-positive bacterial strains as *B. subtilis* and *Rhodococcus sp.* All of the tested compounds showed the good biological activity against diverse bacteria. Results were demonstrated that Zn(II) complex exhibited the highest zones of inhibition i.e. 29±0.4, 26±1.2, 29±0.9, and 28±0.7 mm against *Rhodococcus*, *E. coli*, *B. subtilis* and *P. aeruginosa*, respectively as compared to ligand, standard drug and other complexes. It was





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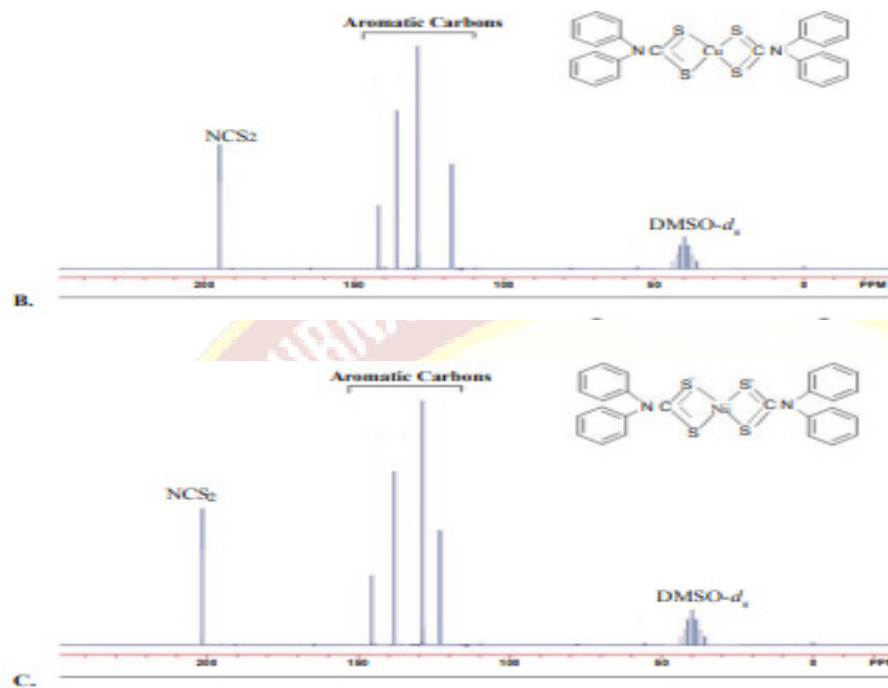
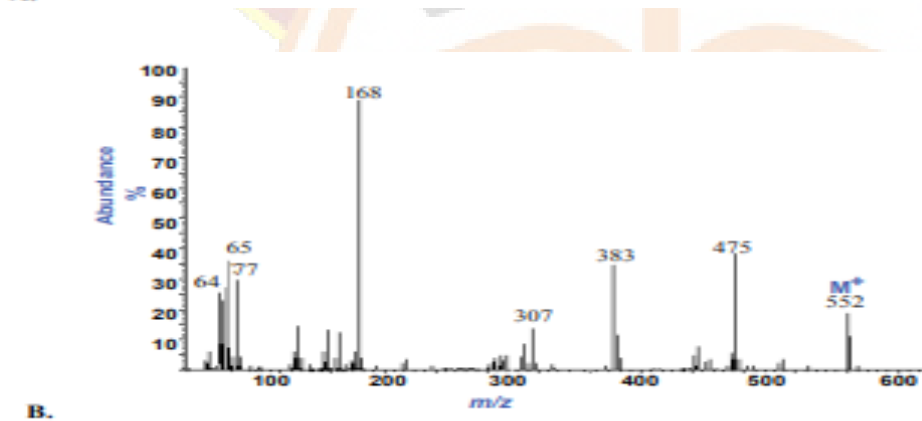
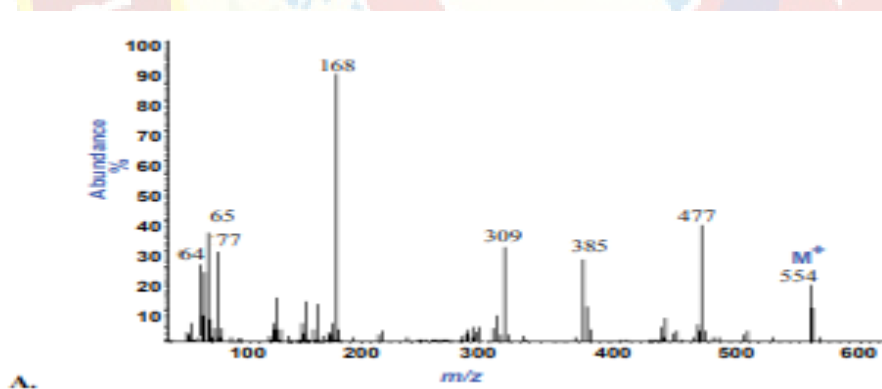


Fig. 3: <sup>13</sup>C NMR spectra of Zn(II)-, Cu(II)- and Ni(II)-DPDTC complexes <sup>13</sup>C NMR spectra of (A) Zn(II)-DPDTC complex; (B) Cu(II)- DPDTC complex and (C) Ni(II)-DPDTC complex





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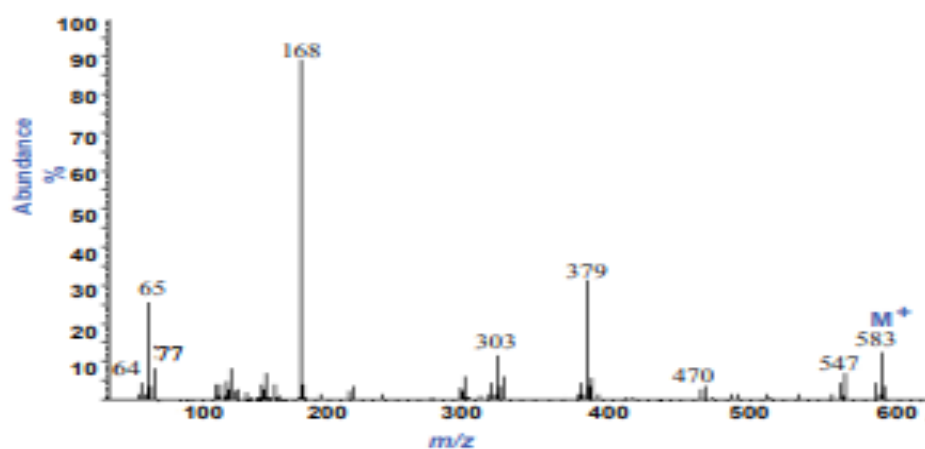


Fig. 4: Mass spectra of Zn(II), Cu(II) and Ni(II)-DPDTC complexes Mass spectra of (A) Zn(II)-DPDTC complex; (B) Cu(II)- DPDTC complex and (C) Ni(II)-DPDTC complex

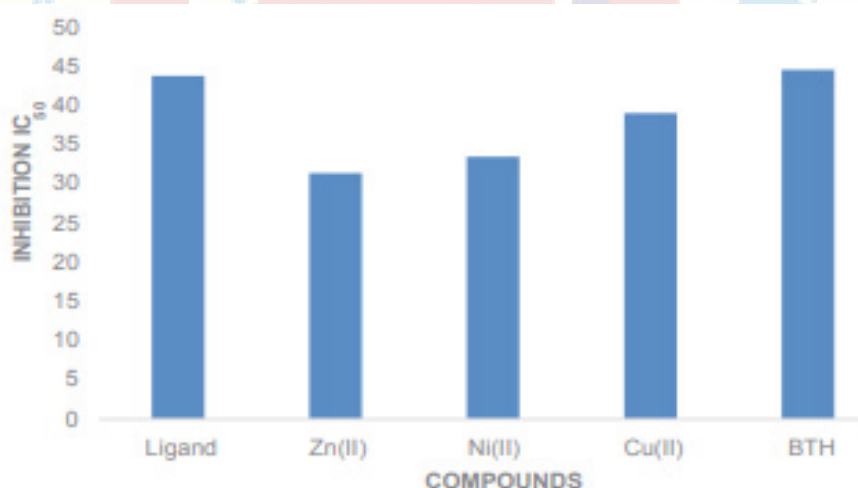


Fig. 5: IC<sub>50</sub> values of antioxidant activity IC<sub>50</sub> inhibition values of DPPH antioxidant activities of ligand and synthesized complexes observed from the results that least antibacterial activity was exhibited by the nickel complex of DPDTC with zones of inhibition 23±0.9, 23±1.6, 21±1.2 and 14±1.4 mm against Rhodococcus, E. coli, B. subtilis and P. aeruginosa, respectively in contrast to zinc and copper complexes[50]. On the other hand, these nickel complexes of DPDTC showed better inhibitory effect on all tested bacterial strains in contrast to ligand as well as the standard drug as shown in Table 1. While Cu(II) complex was found to be more active than the ligand against Rhodococcus and E. coli with zones of inhibition 20±1.6 and 26±0.7 mm, respectively, it also exhibited significant and better antibacterial activity against all tested bacterial strains when compared to the standard drug ampicillin. All results of antibacterial testing revealed that synthesized metal complexes (Zn, Cu and Ni) exhibited greater zones of inhibition against all tested bacterial pathogens when compared to that produced by the standard drug. It was also manifested from the results of antibacterial activity that complexation of ligand with Zn, Cu and Ni could also be responsible for the enhanced antibacterial activity as the results in Table 1 has shown. Many literature reports also revealed the metal complexation of ligand-dependent antibacterial effect and the findings of the current investigation were in agreement with the reported results. The order of antibacterial activity of the compounds was Zn(II)>Cu(II)>Ni(II)>ligand>ampicillin as shown in Table 1. Antifungal activity was tested against A. niger, A. flavus, C. albican and Acetomyceta. It was observed that highest antifungal activity was exhibited by the synthesized Cu(II) complex of DPDTC while lowest antifungal activity was observed with the ligand as shown in Table 2. The synthesized Cu(II) complexes of DPDTC produced larger zone of inhibition of 27±0.8, 21±0.2, 26±0.3 and 29±0.5 mm against A. niger,



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**TABLE 1: ZONES OF INHIBITIONS OF LIGAND AND SYNTHESIZED METAL COMPLEXES AGAINST DIFFERENT BACTERIAL STRAINS**

| Compounds   | Inhibition zone of diameter (mm) |                      |                |                    |
|-------------|----------------------------------|----------------------|----------------|--------------------|
|             | <i>B. subtilus</i>               | <i>P. aeruginosa</i> | <i>E. coli</i> | <i>Rhodococcus</i> |
| Ligand      | 21±0.6                           | 19±0.4               | 16±0.9         | 14±0.5             |
| Zn(II)      | 29±0.9                           | 28±0.7               | 26±1.2         | 29±0.4             |
| Cu(II)      | 23±1.5                           | 19±0.8               | 26±0.7         | 20±1.6             |
| Ni(II)      | 21±1.2                           | 14±1.4               | 23±1.6         | 23±0.9             |
| *Ampicillin | 11±0.4                           | 10±0.6               | 11±0.3         | 12±0.4             |

\*Standard drug; antibacterial species. Zone of Inhibition (mm) include the agar wall diffusion assay hole diameter (4 mm), which carried 100 µg/ml from ligand and their metal complexes solution. The diameter of inhibition zones are means triplicate±SD and p

**TABLE 2: ZONES OF INHIBITIONS OF LIGAND AND SYNTHESIZED METAL COMPLEXES AGAINST DIFFERENT FUNGAL STRAINS**

| Compounds    | Inhibition zone of diameter (mm) |                  |                    |                    |
|--------------|----------------------------------|------------------|--------------------|--------------------|
|              | <i>A. niger</i>                  | <i>A. flavus</i> | <i>C. albicans</i> | <i>Acetomyceta</i> |
| Ligand       | 17±0.4                           | 16±0.7           | 16±0.7             | 18±0.8             |
| Zn(II)       | 24±0.9                           | 19±1.1           | 28±0.4             | 22±0.6             |
| Cu(II)       | 27±0.8                           | 21±0.2           | 26±0.3             | 29±0.5             |
| Ni(II)       | 19±1.5                           | 18±0.3           | 21±0.3             | 23±1.3             |
| *Fluconazole | 12±0.6                           | 15±0.2           | 13±0.1             | 13±0.3             |

\*Standard drug; antifungal species. Zone of Inhibition (mm) include the agar wall diffusion assay hole diameter (4 mm), which carried 100 µg/ml from ligand and their metal complexes solution. The diameter of inhibition zones are means triplicate±SD and p<0.05 is taken as significant)

*A. flavus*, *C. albican* and *Acetomyceta*, respectively in comparison to ligand, fluconazole and other metal complexes (Zn and Ni) tested. The Zn(II) complexes of DPDTC exhibited higher antifungal activity compared to the ligand, the standard drug and Ni complex. It was also observed from the results that among all synthesized metal complexes least inhibitory effect on the growth of *A. niger* (19±1.5 mm), *A. flavus* (18±0.3 mm), *C. albicans* (21±0.3 mm) and *Acetomyceta* (23± 1.3 mm) was produced by the Ni complex of DPDTC, however this inhibitory effect was still greater than that of the ligand and the standard drug as shown in Table 2. The results of antifungal activity of Zn, Cu, Ni complexes of the ligand indicated that the enhancement in antifungal activity is due to complexation with heavy metals[52]. Results also demonstrated that the synthesized metal complexes showed highest antifungal activity against pathogenic fungal strains due also to the presence of DTC group in the complexes[53]. Results of the current study appear to be in agreement with those reported previously[55,56]. Many literature reports also revealed the metal complexation of ligand and DTC group-dependent antifungal effect[57]. The order of antifungal activity of the compounds was Cu(II)>Zn(II)>Ni(II)>ligand>fluconazole. The effect of inhibition of the ligand and synthesized complexes were represented in Table 2. Zn(II), Cu(II) and Ni(II) complexes of DPDTC have been synthesized by the diphenylamine and CS<sub>2</sub> in the presence of NaOH and characterized by spectroscopic analysis. The characterization techniques confirmed that DPDTC ligands were bidentate coordination to the metal atom via sulphur atom of all the synthesized complexes. Conductance measurements indicated the complexes were of ionic nature. The IC<sub>50</sub> of Zn(II) was lower 31.45± 0.31 µM it also showed highest antioxidant activity than ligand and other complexes as well as reference standard (BHT). The order of DPPH scavenging activity of these compounds was Zn(II)>Ni(II)>Cu(II)>BHT>ligand. These metal complexes exhibited antibacterial and antifungal activity against the tested microbial species. The antibacterial activity of the DPDTC ligand and Zn(II) complex indicated the highest antibacterial activity than the other complexes against the tested species in the order of Zn(II)>Cu(II)>DPDTC>Ni(II)>ampicillin. The antifungal activity of Cu(II) complex of DPDTC was found to be the highest against fungal strain as compared to Zn(II) and Ni(II) complexes of DPDTC and the rank order activity was, Cu(II)>Zn(II)>Ni(II)>NaDPDTC>fluconazole. The antimicrobial activity



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might be explained by the interaction of the Cu, Ni, Zn-based complexes with the cytoplasmic membrane of the fungi as well as bacteria. Such interactions with other rich electronic donor centres, amino acids, proteins, nucleosides, carbohydrates and steroids can increase lipid-solubility assisting the complexes to cross the cell membrane. The use of metal-based compounds might represent an alternative therapeutic route to overcome resistance to most used drugs such as fluconazole, amphotericin, among others. In any case, further research is necessary to fully understand and exploit the potential beneficial effects of these complexes.

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